Errata

Issue 14, Winter 2001, page 66, left column.

Do the Polysulfides Really Form as Intermediates?

In an attempt to produce one or more polysulfides, a mixture of potassium carbonate and sulfur were reacted; 0.30 g of [*potassium nitrate* should have been *sulfur*] and 0.20 g of anhydrous potassium carbonate were mixed together. Approximately 0.1 g of this mixture was placed on a metal spatula and heated gently with a propane torch. During the heating process, the melted mixture bubbled (probably CO_2 and trace amounts of other side reaction gases were produced) then the material turned reddish brown and crumbly when cooled. Upon referencing the *Merck Index*^[6] and the *CRC Handbook of Chemistry and Physics*,^[7] it seems most likely that the reddish brown solids were a mixture of K₂S₄ and K₂S₅. Thus, it was concluded that polysulfides form when potassium carbonate is heated with sulfur.